

Exercise 1:

General questions/short exercises:

- a) Are the following loss processes first-order?
 - A. Uptake of CO₂ by the biosphere
 - B. Photolysis of gases in the stratosphere
 - C. Scavenging of aerosol particles by precipitation
- b) Consider a 2-box model with the two boxes being the troposphere (1000-150 hPa) and the stratosphere (150-1 hPa). The lifetime of air in the stratosphere is 2 years. What is roughly the lifetime of air in the troposphere?

Exercise 2:

Briefly comment on following statements: (true, not true, explain why?)

- a). The equilibrium temperature of Venus is lower than that of Earth, even though Venus is nearer to the sun.
- b) Concentrations of CO₂, krypton-85, and other gases emitted mainly in the northern hemisphere DECREASE with altitude in the northern hemisphere but INCREASE with altitude in the southern hemisphere.
- c) The presence of a cloud cover tends to favor lower daytime temperatures and higher nighttime temperatures.

Exercise 3:

Consider a simplified planet-atmosphere system where a thin atmospheric layer is at some distance from the surface of the planet. The albedo of the planet's surface is A. The surface perfectly absorbs infrared radiation; i.e. it can be considered a blackbody in this wavelength region. There is no scattering of sunlight in the atmosphere, so the albedo of the atmosphere by itself is 0. The transmissivity of the atmosphere is τ_s for sunlight and τ_i for infrared radiation. The average incident solar radiation per surface area of the planet is Q. (For clarification: of the incident radiation Q the fraction $\tau_s Q$ is transmitted through the atmosphere, and the fraction $(1-\tau_s)Q$ is absorbed.)

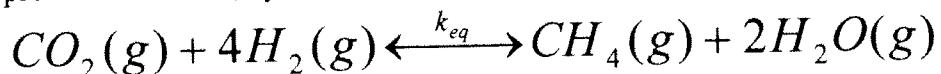
- a) Show that the surface temperature T_0 of the planet is given by following relation:

$$\sigma T_0^4 = Q \left[\frac{1 + \tau_s}{1 + \tau_i} \right] (1 - \tau_s A)$$

- b) Using the relation above show that for some choices of τ_s , τ_i , and A the surface temperature is lower than the equilibrium temperature T_E of the planet. (This would be a sort of negative greenhouse effect).

Exercise 4:

It has been suggested that hydrogen in the Earth's primitive atmosphere led to the production of CH₄ by the reaction



- a) The equilibrium constants k_{eq} for this reaction at 300 and 400 K are 5.2×10^{19} and $2.7 \times 10^{12} \text{ bar}^{-2}$, respectively. If the partial pressures of H₂O, CO₂, and H₂ in the primitive atmosphere were taken to be 3.0×10^{-2} , 3×10^{-4} , and 5.0×10^{-5} bar, respectively, what are the equilibrium pressures of CH₄ at 300 and 400 K?
- b) The equilibrium constants $k_{eq} = k_{\text{forward}}/k_{\text{backward}}$. At 400 K k_{forward} is large, but at 300 K it is immeasurably small. Is it likely that this reaction was responsible for the conversion of much H₂ into CH₄ in the primitive atmosphere? Why, or why not?

Exercise 5:

In the year 2000 a total amount of 3.0×10^{13} kg of fossil fuels has been burned. Assuming that 50% of the emitted CO₂ accumulates in the atmosphere, what is the averaged increase of the CO₂ mixing ratio (in ppm) in the atmosphere? Assume that 80% of the weight of fossil fuels is carbon; the molecular weights of C, O, and air are 12, 16, and 29 g/mol respectively. (Mass of atmosphere: 5.2×10^{18} kg)

Exercise 6 (please on a separate sheet!):

- a) Leg kwalitatief uit wat voor stralingsforcing ontstaat aan het aardoppervlak, in de atmosfeer en aan de top van de atmosfeer als gevolg van aerosol dat zowel zonlicht terugkaatst als absorbeert.
- b) Leg kort uit hoe de aanwezigheid van roet in de atmosfeer het regionale klimaat kan veranderen. Denk hierbij ook aan neerslag.
- c) Beschouw een wolk met een dikte van 200 m, een wolkendruppelconcentratie van 75 cm⁻³ en een druppelstraal van 12.5 μm. Bereken het albedo A , met $A = \tau/(\tau + 6.7)$ en τ de optische dikte.
- d) Nu heeft de wolk uit c) een druppelconcentratie van 225 cm⁻³. Bereken de forcing ΔF ten opzichte van de wolk uit c). De zonneconstante S is 1360 W m⁻².